

## Achievement Standard

**Subject Reference** Chemistry 3.3

**Title** Describe oxidation-reduction processes

**Level** 3                      **Credits** 3                      **Assessment** External

**Subfield** Science

**Domain** Chemistry

**Registration date** 23 November 2005                      **Date version published** 23 November 2005

This achievement standard involves describing oxidation-reduction processes.

### Achievement Criteria

Achievement	Achievement with Merit	Achievement with Excellence
<ul style="list-style-type: none"> <li>Describe oxidation-reduction processes.</li> </ul>	<ul style="list-style-type: none"> <li>Explain and apply oxidation-reduction processes.</li> </ul>	<ul style="list-style-type: none"> <li>Discuss oxidation-reduction processes.</li> </ul>

### Explanatory Notes

- This achievement standard is derived from *Chemistry in the New Zealand Curriculum*, Learning Media, Ministry of Education, 1994, p. 28, achievement objectives 8.1, 8.2 and 8.3.
- Processes* involve reactions and calculations, which may include electrochemical cells and their properties, the use of reduction potentials, and spontaneity of oxidation-reduction reactions.
- Calculations* may include determination of oxidation numbers, mole ratios and those related to electrochemical cells.
- Knowledge of appearance and state of the following reactants and the product to which they are converted in an oxidation-reduction reaction is required.  
Oxidants will be limited to: O<sub>2</sub>, Cl<sub>2</sub>, I<sub>2</sub>, Fe<sup>3+</sup>, dilute acid (with metals), H<sub>2</sub>O<sub>2</sub>, MnO<sub>4</sub><sup>-</sup> (reacting in acidic, basic or neutral conditions), Cu<sup>2+</sup>, Cr<sub>2</sub>O<sub>7</sub><sup>2-</sup>/H<sup>+</sup>, OCl<sup>-</sup>, concentrated HNO<sub>3</sub>, IO<sub>3</sub><sup>-</sup>, MnO<sub>2</sub>.

Reductants will be limited to: metals, C, CO, H<sub>2</sub>, Fe<sup>2+</sup>, Br<sup>-</sup>, I<sup>-</sup>, H<sub>2</sub>S, SO<sub>2</sub>, SO<sub>3</sub><sup>2-</sup>, S<sub>2</sub>O<sub>3</sub><sup>2-</sup>, H<sub>2</sub>O<sub>2</sub>, H<sub>2</sub>C<sub>2</sub>O<sub>4</sub>.  
Appropriate information relating to other oxidants or reductants will be provided.

5 Standard reduction potentials will be included where required.

6 Terms

- *Describe* involves identifying, naming, drawing, giving characteristics of, giving an account of, defining, and/or carrying out simple calculations.
- *Explain and apply* involves describing as well as giving reasons for, making links between chemical concepts and/or observations, or carrying out calculations.
- *Discuss* involves showing understanding by analysing, interpreting, justifying, relating, evaluating, comparing and contrasting, and/or calculating.

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### Quality Assurance

- 1 Providers and Industry Training Organisations must be accredited by the Qualifications Authority before they can register credits from assessment against achievement standards.
- 2 Accredited providers and Industry Training Organisations assessing against achievement standards must engage with the moderation system that applies to those achievement standards.

Accreditation and Moderation Action Plan (AMAP) reference

0226